Synthesis and characterization of ceria nanomaterials.

Nitzia Cheong Ng

University of Louisville

Follow this and additional works at: https://ir.library.louisville.edu/etd

Recommended Citation
https://doi.org/10.18297/etd/1054

This Master's Thesis is brought to you for free and open access by ThinkIR: The University of Louisville's Institutional Repository. It has been accepted for inclusion in Electronic Theses and Dissertations by an authorized administrator of ThinkIR: The University of Louisville's Institutional Repository. This title appears here courtesy of the author, who has retained all other copyrights. For more information, please contact thinkir@louisville.edu.
SYNTHESIS AND CHARACTERIZATION OF CERIA NANOMATERIALS

By

Nitzia Cheong Ng
B.S., University of Louisville, May 2009

A Thesis
Submitted to the Faculty of the
University of Louisville
J. B. Speed School of Engineering
as Partial Fulfillment of the Requirements
for the Professional Degree

MASTER OF ENGINEERING

Department of Chemical Engineering

August 2010
SYNTHESIS AND CHARACTERIZATION OF CERIA NANOMATERIALS

Submitted by: __________________________________
Nitzia Cheong Ng

A Thesis Approved On

___________________________________
Date

by the Following Reading and Examination Committee:

___________________________________
Dr. Xiao-An Fu, Thesis Advisor

___________________________________
Dr. James C. Watters

___________________________________
Dr. Gamini Sumanasekera
ACKNOWLEDGEMENTS

The author is grateful for the support, knowledge, and patience of Dr. Xiao-An Fu, Thesis Advisor.

The author thanks Jacek Jasinski and Rodica McCoy for their help in characterization.

The author appreciates Dr. James Watters for participating in the defense committee.

The author appreciates Dr. Gamini Sumanasekera for participating in the defense committee.

The author also thanks Surender Venna, Chinmay Deshmane and Dr. Moises Carreon for their help and knowledge in BET and Dr. Kyung Kang and Dr. Martin O’Toole for the centrifugation instrument.

The author is gratified to Dr. Michael Tseng from the Medical School for his knowledge and support.
ABSTRACT

Cerium dioxide or ceria, CeO$_2$, has been widely used in industry as catalyst for automotive exhaust controls, chemical mechanical polishing (CMP) slurries, and high temperature fuel cells because of its unique metal oxide properties. This well-known rare metal oxide has high thermal stability, electrical conductivity and chemical diffusivity. Proper synthesis method requires knowledge of reaction temperature, concentration, and time effects on the synthesis. In this work, ceria nanomaterials were prepared via the hydrothermal method using a Teflon autoclave. Cerium nitrate solution was used as the source and three different precursors: NaOH, H$_2$O$_2$, and NH$_4$OH were used as the oxidizing agents. CeO$_2$ nanoplates, nanocubes and nanorods were produced and studied using transmission electron microscopy (TEM), BET specific surface area, X-ray diffraction (XRD) and X-ray photoelectron spectroscopy (XPS).

Through characterization, CeO$_2$ nanomaterials showed the presence of mixed valence states (Ce$^{3+}$ and Ce$^{4+}$) through XPS spectra. Deconvolution was performed to investigate the ratio of Ce$^{3+}$/Ce$^{4+}$ concentration in the synthesized CeO$_2$ nanostructures. Nanocubes showed a higher Ce$^{3+}$ concentration. CeO$_2$ nanomaterials were found to be mesoporous. Nanoplates synthesized with H$_2$O$_2$, and NH$_4$OH were found with surface areas of 95.11 m$^2$/ g and 62.07 m$^2$/ g, respectively. Nanorods and nanocubes showed surface areas of 16.77 m$^2$/ g and 16.55 m$^2$/ g, respectively. The prepared ceria nanoplates, nanocubes and nanorods had crystallite size in the range of 5-25 nm and pore size range of 7-15 nm. XRD spectra confirmed that the peaks were indexed to the cubic phase of CeO$_2$. 
CeO$_2$ with fluorite structure and with an average lattice parameter, 5.407 Å. Higher Ce$^{3+}$ concentration and exposed surface of crystalline planes suggest that nanorods are better catalyst for CO oxidation and oxygen storage capacity (OSC).
# TABLE OF CONTENTS

<table>
<thead>
<tr>
<th>Section</th>
<th>Page</th>
</tr>
</thead>
<tbody>
<tr>
<td>APPROVAL PAGE</td>
<td>ii</td>
</tr>
<tr>
<td>ACKNOWLEDGEMENTS</td>
<td>iii</td>
</tr>
<tr>
<td>ABSTRACT</td>
<td>iv</td>
</tr>
<tr>
<td>TABLE OF CONTENTS</td>
<td>vi</td>
</tr>
<tr>
<td>NOMENCLATURE</td>
<td>viii</td>
</tr>
<tr>
<td>LIST OF TABLES</td>
<td>x</td>
</tr>
<tr>
<td>LIST OF FIGURES</td>
<td>xi</td>
</tr>
<tr>
<td>I. INTRODUCTION</td>
<td>1</td>
</tr>
<tr>
<td>II. BACKGROUND</td>
<td>3</td>
</tr>
<tr>
<td>A. Cerium dioxide or Ceria</td>
<td>3</td>
</tr>
<tr>
<td>B. Hydrogen Peroxide</td>
<td>6</td>
</tr>
<tr>
<td>C. Ammonium Hydroxide</td>
<td>8</td>
</tr>
<tr>
<td>D. Sodium Hydroxide</td>
<td>9</td>
</tr>
<tr>
<td>III. EXPERIMENTATION</td>
<td>13</td>
</tr>
<tr>
<td>A. Plan of Experimentation</td>
<td>13</td>
</tr>
<tr>
<td>B. Materials</td>
<td>14</td>
</tr>
<tr>
<td>C. Procedure</td>
<td>15</td>
</tr>
<tr>
<td>D. Characterization of the Prepared CeO$_2$</td>
<td>17</td>
</tr>
<tr>
<td>E. Equipment</td>
<td>17</td>
</tr>
<tr>
<td>IV. RESULTS AND DISCUSSION</td>
<td>20</td>
</tr>
<tr>
<td>A. Shape Selective and Mechanism</td>
<td>20</td>
</tr>
<tr>
<td>B. TEM analysis</td>
<td>21</td>
</tr>
<tr>
<td>C. BET specific surface area analysis</td>
<td>26</td>
</tr>
<tr>
<td>D. XPS analysis</td>
<td>31</td>
</tr>
<tr>
<td>E. XRD analysis</td>
<td>37</td>
</tr>
<tr>
<td>V. CONCLUSIONS</td>
<td>48</td>
</tr>
<tr>
<td>VI. RECOMMENDATIONS</td>
<td>49</td>
</tr>
<tr>
<td>APPENDIX 1: RAW DATA</td>
<td>50</td>
</tr>
<tr>
<td>Section</td>
<td>Page</td>
</tr>
<tr>
<td>---------------------------------</td>
<td>------</td>
</tr>
<tr>
<td>APPENDIX 2: DECONVOLUTED XPS SPECTRUM</td>
<td>52</td>
</tr>
<tr>
<td>APPENDIX 3: SAMPLE CALCULATION</td>
<td>54</td>
</tr>
<tr>
<td>REFERENCES</td>
<td>55</td>
</tr>
<tr>
<td>VITA</td>
<td>59</td>
</tr>
</tbody>
</table>
NOMENCLATURE

\begin{align*}
\text{NH}_4\text{OH} & = \text{Ammonium hydroxide} \\
\text{BE} & = \text{Binding energy} \\
\text{BJH} & = \text{Barret-Joyner-Halenda} \\
\text{BET} & = \text{Brunaure Emmett Teller} \\
\text{CO} & = \text{Carbon dioxide oxidation} \\
\text{CeO}_2 & = \text{Ceria or cerium oxide} \\
\text{Ce(NO}_3\text{)}_3 & = \text{Cerium nitrate} \\
\text{C}_{\text{NaOH}} & = \text{Concentration of sodium hydroxide} \\
\text{d}_{\text{hkl}} & = \text{Crystalline face space at the crystallite indices} \\
\text{D}_{\text{hkl}} & = \text{Crystallite size at the crystallite indices} \\
\text{h},\text{k},\text{l} & = \text{Crystallite size indices} \\
\text{TG-DTA} & = \text{Differential thermal analysis and thermo-gravimetric analysis} \\
\theta & = \text{Diffraction angle} \\
\text{FCC} & = \text{Face–centered-cubic} \\
\text{FFT} & = \text{Fast Fourier transform} \\
\text{FT-IR} & = \text{Fourier transform infrared} \\
\text{H}_2\text{O}_2 & = \text{Hydrogen peroxide} \\
\text{HRTEM} & = \text{High-resolution transmission electron microscopy} \\
n & = \text{Integer number, usually 1} \\
A_i & = \text{Integrated area of peak “i.”} \\
a & = \text{Lattice parameter} \\
\text{B}_{\text{hkl}} & = \text{Measured half-width}
\end{align*}
<table>
<thead>
<tr>
<th>Abbreviation</th>
<th>Definition</th>
</tr>
</thead>
<tbody>
<tr>
<td>OSC</td>
<td>Oxygen storage capacity</td>
</tr>
<tr>
<td>SAED</td>
<td>Selected area electron diffraction</td>
</tr>
<tr>
<td>NaOH</td>
<td>Sodium hydroxide</td>
</tr>
<tr>
<td>H₂-TPR</td>
<td>Temperature-programmed reduction of hydrogen</td>
</tr>
<tr>
<td>TEM</td>
<td>Transmission electron microscope</td>
</tr>
<tr>
<td>λ</td>
<td>Wavelength of the incident X-rays</td>
</tr>
<tr>
<td>XPS</td>
<td>X-ray photoelectron spectroscopy</td>
</tr>
<tr>
<td>XRD</td>
<td>X-ray diffraction</td>
</tr>
</tbody>
</table>
LIST OF TABLES

TABLE II-I CRYSTAL STRUCTURES, SHAPES, AND SIZES OF SEVERAL CEO₂ SAMPLES 10

TABLE III-I SYNTHESIS OF CERIA NANOPLATE WITH HYDROGEN PEROXIDE PRECURSOR 15

TABLE III-II SYNTHESIS OF CERIA NANOPLATE WITH AMMONIUM HYDROXIDE PRECURSOR 15

TABLE III-III SYNTHESIS OF CERIA NANOROD (C NaOH=9M) WITH SODIUM HYDROXIDE PRECURSOR 16

TABLE III-IV SYNTHESIS OF CERIA NANOCUBE (C NaOH=6M) WITH SODIUM HYDROXIDE PRECURSOR 16

TABLE IV-I BET SPECIFIC SURFACE AREA OF DIFFERENT CEO₂ SHAPES 27

TABLE IV-II PORE SIZE AND PORE VOLUME OF THE PREPARED CERIA MATERIALS 30

TABLE IV-III XPS BINDING ENERGIES OF INDIVIDUAL PEAKS OF THE CE (3D) SPECTRUM FOR DIFFERENT CeO₂ SHAPES 35

TABLE IV-VI INTEGRATED AREAS OF INDIVIDUAL PEAKS 36

TABLE IV-V CONCENTRATION OF Ce³⁺ FOR DIFFERENT SHAPES OF CERIA CALCULATED USING EQ. (1). 36

TABLE IV-VI COMPARISON OF LATTICE PARAMETERS AT DIFFERENT CALCINATION TEMPERATURES 41

TABLE IV-VII LATTICE TYPE AND CRYSTALLITE SIZE OF CeO₂ NANOSTRUCTURES 42

TABLE IV-VIII BET SPECIFIC SURFACE AREA AND CRYSTALLITE SIZE OF DIFFERENT CeO₂ SHAPES 44

TABLE A1 – 1 EXPERIMENTAL DATA FOR CERIA NANOPLATES BY H₂O₂ 50

TABLE A1 – 2 EXPERIMENTAL DATA FOR CERIA NANOPLATES BY NH₄OH 50

TABLE A1 – 3 EXPERIMENTAL DATA FOR CERIA NANORODS AND NANOCUBES 51
LIST OF FIGURES

FIGURE 2.1 - Number of publications on CeO2-based materials in catalysis: (■) open literature, (●) patents (Trovarelli, 1999).

FIGURE 2.2- Number of scientific articles (not patents) published in 1997 on ceria and related materials in different areas of catalysis: 1 (TWC), 2 (flue gas treatment), 3 (oxidation), 4 (treatment of diesel exhaust), 5 (catalyst characterization), 6 (electrocatalysis), 7 (hydrogenation), 8 (others) (Trovarelli, 1999).

FIGURE 2.3 - a) TEM image of CeO2 nanowires, b) HRTEM image of individual CeO2 nanowire (Tang, 2005).

FIGURE 2.4 – HRTEM image of CeO2 nanowires (Han, 2005).

FIGURE 2.5 - TEM images of the CeO2 nanoparticles (a and b), nanorods (c and d) and nanowires (e and f) (Tana, 2009).

FIGURE 2.6 - TEM images of the products obtained by the utilization of varied cerium source: (A) (NH4)2Ce(NO3)6; (B) CeCl3·7H2O (Yang, 2009).

FIGURE 3.1 - Experimental flow chart of synthesis of CeO2 with H2O2.

FIGURE 3.2 – Vulcan Furnace

FIGURE 3.3 – Hydrothermal Autoclave

FIGURE 3.4 - Transmission Electron Microscope

FIGURE 3.5 - X-ray Photoelectron Spectroscopy

FIGURE 3.6 - X-ray Diffraction

FIGURE 4.1 - TEM and SAED images of CeO2 nanoplates at 220°C by H2O2.

FIGURE 4.2 - TEM, HRTEM and SAED images of CeO2 nanoplates by H2O2 synthesized at different temperatures (a) 180°C, (b) 200°C, and (c) 250°C for 3h.

FIGURE 4.3 - TEM and SAED image of CeO2 nanoplates by NH4OH at 100°C.

FIGURE 4.4- TEM and HRTEM images of CeO2 nanorods with a CNaOH= 9 M at 100°C.
FIGURE 4.5 - TEM and HRTEM images of CeO₂ nanocubes with a $C_{NaOH} = 9$ M at 180°C.

FIGURE 4.6 - TEM and HRTEM images of CeO₂ nanocubes with a $C_{NaOH} = 6$ M at 180°C.

FIGURE 4.7 - Nitrogen adsorption-desorption isotherm pattern as synthesized of CeO₂ nanoplates by H₂O₂ (a), nanoplates by NH₄OH (b), nanocubes (c), and nanorods (d). (● adsorption, ■ desorption).

FIGURE 4.8 - Plot of the pore volume vs. pore diameter for CeO₂ nanoplates by H₂O₂, nanoplates by NH₄OH, nanocubes, and nanorods in BJH adsorption (a) and desorption (b).

FIGURE 4.9 - XPS survey spectra of CeO₂ nanoplates by H₂O₂ and NH₄OH, nanocubes, and nanorods.

FIGURE 4.10 - XPS spectrum of the six peaks of Ce 3d for nanoplates by H₂O₂ (a), nanorods (b), nanoplates by NH₄OH (c), and nanocubes (d).

FIGURE 4.11 - XPS spectrum of the synthesized ceria nanoplates by H₂O₂ with the mixed valence state.

FIGURE 4.12 - Deconvoluted XPS Ce (3d) spectrum of ceria nanocubes.

FIGURE 4.13 - XRD of CeO₂ nanorods (a), nanoplates by NH₄OH (b), nanocubes (c), and nanoplates by H₂O₂ (d).

FIGURE 4.14 - XRD patterns of CeO₂ nanoplates by H₂O₂ calcined at 100°C (a) and 400°C (b).

FIGURE 4.15 - XRD patterns of CeO₂ nanoplates by NH₄OH calcined at 100°C (a) and 600°C (b).

FIGURE 4.16 - XRD patterns of CeO₂ nanocubes calcined at 100°C (a) and 400°C (b).

FIGURE 4.17 - XRD patterns of CeO₂ nanorods calcined at 100°C (a) and 600°C (b).

Figure 4.18 - Crystallite size as a function of calcination temperatures for nanoplates by H₂O₂ (a), nanocubes (b), nanoplates by NH₄OH (c), and nanorods (d).
FIGURE 4.19 – Concentration of Ce$^{3+}$ as a function of lattice parameter for CeO$_2$ nanocubes (a), nanorods (b), nanoplates by NH$_4$OH (c), and nanoplates by H$_2$O$_2$.

FIGURE A2-1 – CeO$_2$ nanoplates by H$_2$O$_2$ deconvoluted XPS spectrum

FIGURE A2-2 – CeO$_2$ nanoplates by NH$_4$OH deconvoluted XPS spectrum

FIGURE A2-3 – CeO$_2$ nanorods deconvoluted XPS spectrum

FIGURE A2-4 - CeO$_2$ nanocubes deconvoluted XPS spectrum
I. INTRODUCTION

Studies in the morphologies and sizes of cerium dioxide or ceria have attracted the interests of researchers in recent years due to ceria’s unique properties that lead to a wide range of applications. Some examples of ceria applications include fast ion conductors, UV blockers, catalyst (Tang, 2005), gas and bio-sensors, cosmetic material, fuel cells, and slurry for chemical mechanical planarization (Seok, 2004). Several synthetic approaches such as electrochemical deposition, microwave-hydrothermal, solvothermal, hydrothermal and others have been developed to form ceria nanomaterials. Of these methods, the hydrothermal method (Mai, 2005) has been considered as one of the most effective and economical routes, because of its single step low-temperature synthesis, superior composition and morphological control, and powder reactivity.

To explore the morphology and properties of ceria, proper synthesis and precursors are required to obtain specific ceria nanostructures. The present work investigates the different shapes of ceria formed by using three different precursors: hydrogen peroxide, ammonium hydroxide, and sodium hydroxide through hydrothermal
method for 24 hours. The different precursors and reaction parameters such as time, temperature, and concentration influence the size and shape of ceria, forming nanoplates, nanorods, and nanocubes.

In the current project, ceria is synthesized from cerium nitrate and three different precursors mentioned earlier. Once the synthesis was done, the samples were centrifuged, purified by washing several times, and dried. The dried precipitates show the following colors: white for the synthesis with H$_2$O$_2$, light yellow with NaOH, and medium yellow with NH$_4$OH. Characterization is performed to understand the morphologies and properties of ceria formed using each precursor. In this work, the samples have been characterized with TEM, X-Ray diffraction, BET specific surface area, and XPS to understand their properties.
II. BACKGROUND

A. Cerium Oxide or Ceria

Cerium with an atomic number 58 has brought new discoveries in terms of nanomaterials. Cerium is the most abundant of the rare earth elements. As a well-known functional rare earth material, ceria has wide applications in fields such as catalysis, electrochemistry, and optics due to its unique properties (Powell, 1988). In addition, cerium dioxide has been extensively studied and employed in various applications including fast ion conductors, oxygen storage capacitors, catalysts, UV blockers, polishing materials (Tang, 2005), and gas sensor (Yahiro, 1988).

Preparation of ceria powders is by no means a new research subject; however, practical methods are still needed for synthesizing high-quality ultrafine powders with required characteristics in terms of their size, uniformity, morphology, specific surface area, and crystallinity. Numerous techniques have been proposed to synthesize nano-sized CeO$_2$ particles with promising control of properties (Lee, 2004; Zhou, 1995).
Cerium dioxide has been used in many catalysis applications. Trovarelli (1996 and 2002) reported that CeO$_2$ has been extensively studied and applied in heterogeneous catalysis based on its ability to release and uptake oxygen under operation conditions while preserving the fluorite crystal structure. It has an important role in two of the most important commercial catalytic processes in terms of economic relevance and tonnage: three-way catalysis (TWC) and fluid catalytic cracking (FCC). The importance of CeO$_2$ in catalysis is also exhibited by the number of industrial and academic publications that have appeared on the topic over the last several years as shown in Figure 2.1. In 1997, the number of published ceria topics reveals that approximately 50% is based on the treatment of emissions from mobile and stationary sources, and that the majority of articles deal with the application of ceria in TWCs and novel technologies for the treatment of emissions from diesel and spark-ignited internal combustion engines as shown in Figure 2.2 (Trovarelli, 1999).

![FIGURE 2.1 - Number of publications on CeO2-based materials in catalysis: (■) open literature, (●) patents (Trovarelli, 1999).]
This indicates that there are several rising applications or processes for which cerium oxide is currently being actively investigated. Trovarelli mentioned that ceria acts as an oxygen buffer by absorbing and releasing oxygen through a fast Ce$^{3+}$/Ce$^{4+}$ cycle in the TWCs, and simultaneously encouraging the oxidation of CO to CO$_2$ and involving the participation of the lattice oxygen species.

![Graph showing the number of scientific articles published in 1997 on ceria and related materials in different areas of catalysis.](image)

**FIGURE 2.2**: Number of scientific articles (not patents) published in 1997 on ceria and related materials in different areas of catalysis: 1 (TWC), 2 (flue gas treatment), 3 (oxidation), 4 (treatment of diesel exhaust), 5 (catalyst characterization), 6 (electrocatalysis), 7 (hydrogenation), 8 (others) (Trovarelli, 1999).

Additionally, the unique redox property is usually described by the oxygen storage capacity (OSC), which is largely dependent on the size of ceria particles (Carrettin, 2004; Trovarelli, 1996 and 2002). The morphology also contributes to the redox feature of ceria depending on the exposed ceria crystalline planes (Mai; 2005; Zhou, 2005). Lastly, it was found that CeO$_2$ has potential uses (Trovarelli, 1996) as diesel
fuels catalyst for a more complete combustion to abate soot formation, and for higher-energy efficiency and energy density of solid-oxide fuel-cell.

B. Hydrogen Peroxide

The preparation of ceria using hydrogen peroxide as an oxidizer has been a facile method through hydrothermal synthesis. Several methods have been used to synthesize ceria nanowires by cerium nitrate (Ce(NO$_3$)$_3$) with H$_2$O$_2$. For example, Figure 2.3 shows a TEM image of CeO$_2$ nanowires by hydrothermal method. The obtained CeO$_2$ nanowires are single crystalline and uniform with diameters of 20-70 nm and lengths up to 40 μm (Tang, 2005).

![Figure 2.3 - a) TEM image of CeO$_2$ nanowires, b) HRTEM image of individual CeO$_2$ nanowire (Tang, 2005).](image)

On the other hand, H$_2$O$_2$ with cerium nitrate can form CeO$_2$ nanoplates as described in Chapter IV. The morphology (Mai, 2005; Zhou, 2005) is mainly composed
of polyhedron which exposes eight \{1 1 1\} or eight \{1 1 1\} and six \{1 0 0\} plane, but the rod-shaped nanostructures tend to preferentially expose four \{1 1 0\} and two \{1 0 0\} planes. Therefore, Mai et al. (2005) reported that the oxygen storage is confined mainly on the surface of \(\text{CeO}_2\) nanoparticles, while the oxygen species both on the surface and in the bulk could participate in the redox cycle in the nanorods, giving a much higher OSC (Mai, 2005). Djuricic et al. (1999) reported that the role of \(\text{H}_2\text{O}_2\) is to make it much easier to change the valence state of cerium ion from Ce (III) to Ce (IV). The following reaction takes place in solution:

\[
\text{Ce}^{3+} + \text{H}_2\text{O}_2 + 2\text{H}^+(\text{aq.}) \rightarrow \text{Ce}^{4+} + 2\text{H}_2\text{O}
\]

To obtain cerium(IV) oxide from Ce(III) nitrate, \(\text{H}_2\text{O}_2\) is used as an oxidizer, owing to the H-bond interaction with other hydroxyl or hydroperoxyl groups at two-directions. Tang et al. (2005) revealed that \(\text{H}_2\text{O}_2\) was also a template agent. In contrast with most of the other templates, which can generate complicated process and an increase of impurity concentration in the final product, \(\text{H}_2\text{O}_2\) makes the process very simple and does not leave any impurity in the hydrothermal reaction system. The \(\text{H}_2\text{O}_2\)-assisted hydrothermal method might offer an opportunity to synthesize other similar nanomaterials and have significant contribution for industrial applications. \(\text{CeO}_2\) nanoparticles have significant impact in applications such as UV blocking and shielding materials. This is because of their considerably small sizes, high surface areas, and especially the general improvements in the material properties with respect to their micronized or bulk like materials (Mai, 2005).
C. Ammonium Hydroxide

Ceria nanoplates are also formed by using the second precursor ammonium hydroxide. Recently, nanoplates have attracted increasing attention due to two advantages: one is high crystallinity and well-defined chemical composition as well as extremely high anisotropy with an ultrathin thickness (Zhang, 2005) and the other is that they are superior precursors to be used for conversion into other nanostructures (Tian, 2002), such as nanotubes based on the rolling-up mechanism. Therefore, CeO$_2$ nanoplates are hoped to be synthesized so that they can improve the catalytic performance and also be easy to convert into other useful nanostructures.

Besides producing CeO$_2$ nanoplates, other approaches have successfully formed CeO$_2$ nanoparticles by adding aqueous ammonium hydroxide precipitant into cerium nitrate solution at room temperature and introducing oxygen into the reactor to oxidize Ce$^{3+}$ to Ce$^{4+}$ (Zhou, 2002). Han et al (2005) proved that CeO$_2$ nanowires with a diameter range from 5-30 nm and a length up to several microns shown in Figure 2.4 can be formed by adding aqueous ammonium hydroxide precipitant into cerium nitrate at 70°C and then aging at 0 °C for 1 day. Chen et al. (2004) reported that high-alkaline environments support the oxidation of Ce(OH)$_3$ to hydrated Ce (IV). Oxidation of Ce$^{3+}$ to Ce$^{4+}$ in solution takes place at high pH, for example, Ce$^{3+}$ + H$_2$O $\rightarrow$ Ce (OH)$^{3+}$ + H$^+$ + e$^-$ with subsequent hydrolysis to Ce(OH)$_4$ and precipitation. Hydroxyl ions play an important role in this process and strongly affect the upper saturation degree of initial precipitate and oxidation of Ce (III) to Ce (IV).
An enhanced catalytic activity was found for CO oxidation by using CeO$_2$ nanoplates as compared with CeO$_2$ nanotubes and nanorods, and the crystal surfaces (100) of CeO$_2$ nanoplates were considered to play an important role in determining their catalytic oxidation properties. The structural characteristics of CeO$_2$ are believed to affect their performance in catalysis and can be extended to other technological applications (Pan, 2008).

D. Sodium Hydroxide

In order to achieve ceria nanorods and nanocubes, sodium hydroxide precursor helps to obtain these morphologies. Cerium nitrate is synthesized with sodium hydroxide to form different crystal structures, shapes and sizes. Table I presents the synthesis factors used to obtain these types of CeO$_2$ morphologies. Mai et al. (2005) found that CeO$_2$ nanorods and nanocubes could be selectively obtained by varying the NaOH concentration and hydrothermal temperature. For example, CeO$_2$ nanorods were
yielded at a much higher concentration \( (C_{\text{NaOH}} = 6 – 9 \text{ mol L}^{-1}) \) when hydrothermally treated at 100°C, while CeO\(_2\) nanocubes were formed under the same reactant concentrations but 180°C. The base concentration and hydrothermal temperature seemed to be the two key factors in the selective formation of CeO\(_2\) nanopolyhedra, nanorods, and nanocubes according to Table II-I.

### TABLE II-I

**CRYSTAL STRUCTURES, SHAPES, AND SIZES OF SEVERAL CeO\(_2\) SAMPLES**

(MAI, 2005)

<table>
<thead>
<tr>
<th>no.</th>
<th>(C_{\text{NaOH}}) (mol L(^{-1}))</th>
<th>(T) (°C)</th>
<th>(t) (h)</th>
<th>structure</th>
<th>shape</th>
<th>size (nm)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>0.01</td>
<td>100</td>
<td>24</td>
<td>cubic</td>
<td>polyhedra</td>
<td>11.5 ± 1.8</td>
</tr>
<tr>
<td>2</td>
<td>0.01</td>
<td>180</td>
<td>24</td>
<td>cubic</td>
<td>polyhedra</td>
<td>9–25</td>
</tr>
<tr>
<td>3</td>
<td>1</td>
<td>100</td>
<td>24</td>
<td>cubic</td>
<td>polyhedra; rods</td>
<td></td>
</tr>
<tr>
<td>4</td>
<td>3</td>
<td>100</td>
<td>24</td>
<td>cubic</td>
<td>polyhedra; rods</td>
<td></td>
</tr>
<tr>
<td>5</td>
<td>6</td>
<td>100</td>
<td>24</td>
<td>cubic</td>
<td>rods</td>
<td>(9.6 ± 1.2) (\times) (50–200)</td>
</tr>
<tr>
<td>6</td>
<td>6</td>
<td>140</td>
<td>24</td>
<td>cubic</td>
<td>rods; cubes</td>
<td></td>
</tr>
<tr>
<td>7</td>
<td>6</td>
<td>180</td>
<td>24</td>
<td>cubic</td>
<td>cubes</td>
<td>36.1 ± 7.1</td>
</tr>
<tr>
<td>8</td>
<td>9</td>
<td>100</td>
<td>24</td>
<td>cubic</td>
<td>rods</td>
<td>(13.3 ± 2.8) (\times) (100–400)</td>
</tr>
</tbody>
</table>

On the other hand, CeO\(_2\) nanowires (Tana, 2009) exhibited a much higher OSC than the nanorods despite their comparable specific surface area and similar geometrical configuration. The higher OSC might be due to the differing proportions of exposed (1 1 0) and/or (1 0 0) planes between the nanorods and the nanowires. Since the OSC on the (1 0 0) plane is much higher than that on the (1 1 0) plane in CeO\(_2\), it is most likely that the nanowire exposes relatively more (1 0 0) planes on the surface than the nanorod, although both of them preferentially expose the (1 1 0) and (1 0 0) planes, as shown in Figure 2.5 (c) to (f).

CeO\(_2\) nanocubes are formed not only due to the accelerated dissolution/recrystallization, but also due to a transfer step from Ce(OH)\(_3\) to CeO\(_2\).
Nanocubes resulted at higher base concentrations as temperature was raised from 100 to 180°C. It is enclosed by six (100) planes and that {100}/ {110}-dominated surface structures are much more reactive for CO oxidization and make the associated lattice oxygen migration from bulk to surface much easier than the {111}-dominated one, and the {100}-dominated one is more active than the {110}-dominated one (Madier, 1999).

FIGURE 2.5 - TEM images of the CeO2 nanoparticles (a and b), nanorods (c and d) and nanowires (e and f) (Tana, 2009).
Other approaches or recent studies in CeO$_2$ have focused on the size/shape-controlled nanostructures. For example, Chen et al. (2004) fabricated polycrystalline CeO$_2$ nanowires via a solution-phase route using sodium bis(2-ethylhexyl) sulfosuccinate as a structure-directing agent, and observed size-dependent effect on Raman spectra. Wu et al. (2004) obtained fluorite-type ceria nanowires with diameters of 70 nm by a sol-gel process within the nanochannels of porous anodic alumina (a “hard” template). Li et al. (2005) synthesized single-crystalline CeO$_2$ nanorods by hydrothermal method, which showed an enhanced CO conversion activity. Yang et al. (2006) produced CeO$_2$ nanocubes through a solvothermal approach using oleic acid as stabilizing agent while monodisperse CeO$_2$ nanocubes have been fabricated via an acrylamide-assisted hydrothermal route by using (NH$_4$)$_2$Ce(NO$_3$)$_6$ and CeCl$_3$·7H$_2$O instead of Ce(NO$_3$)$_3$·6H$_2$O as the cerium source which resulted in uniform spherical particles shown in Figure 2.6 with an average diameter of 200 nm.

![Figure 2.6](image)

**FIGURE 2.6** - TEM images of the products obtained by the utilization of varied cerium source: (A) (NH$_4$)$_2$Ce(NO$_3$)$_6$; (B) CeCl$_3$·7H$_2$O (Yang, 2009).
III. EXPERIMENTATION

A. Plan of Experimentation

Three different synthesis methods were designed to investigate the properties of cerium oxide nanostructures synthesized with the three precursors H₂O₂, NH₄OH, and NaOH. The synthesis was performed by hydrothermal reaction at temperatures 100, 180, and 220°C. The experimental flow chart in Figure 3.1 illustrates the schematic procedure of the synthesis of CeO₂ by hydrothermal method using H₂O₂. The details of the different parameters used to synthesize the different ceria nanostructures are presented in the Procedure Section of this chapter.

FIGURE 3.1 - Experimental flow chart of synthesis of CeO₂ with H₂O₂.
In order to identify the morphology of ceria, characterization was performed through BET specific surface area, transmission electron microscopy (TEM), X-ray photoelectron spectroscopy (XPS), and X-ray diffraction (XRD). TEM, XPS and XRD equipments are illustrated in Figures 3.3 – 3.5. It was found that the synthesis of ceria produces nanoplates, nanorods, and nanocubes depending on the type of precursor used and other parameters such as reaction time and concentration.

B. Materials

Source:
Cerium (III) nitrate hexahydrate, 99% metal basis (0.05 M Ce (NO\textsubscript{3})\textsubscript{3})
Sigma-Aldrich, Inc
St Louis, MO 63103 USA
Product of China
Cat. No.: 238538-100G

Precursors:
Perdrogen® 30% by weight H\textsubscript{2}O\textsubscript{2} puriss. p.a.
Sigma-Aldrich, Inc
St Louis, MO 63178 USA
Made in Germany
Cat. No.: 31642

Sodium Hydroxide, pellets, 99.99% metals basis, semiconductor grade (Na\textsubscript{2}CO\textsubscript{3} < 1%)
Sigma-Aldrich, Inc
St Louis, MO 63103 USA
Product of USA
Cat. No.: 306576-100G

Ammonium Hydroxide, A.C.S. reagent (28%)
Sigma-Aldrich, Inc
St Louis, MO 63103 USA
Product of USA
Cat. No.: 306576-100G

Deionized water
C. Procedure

The hydrothermal synthesis of cerium oxide was conducted in a 50 ml Teflon autoclave shown in Figure 3.3 and then heated up to 100, 180, and 220°C in a furnace, Figure 3.3, for 24 hours. First, a 50 ml solution of 0.05 M Ce(NO₃)₃ was prepared. The weight of 0.05 M Ce(NO₃)₃ for the 50 ml solution was 1.085 g. This solution was used to react with H₂O₂ or NH₄OH to form ceria nanoplates.

Table III-I through table III-IV present the three methods of synthesis that were performed; they differ because of the precursors used for the formation. In this study, ceria nanoplates, nanorods, and nanocubes are produced by the following the synthesis recipes.

**TABLE III-I**

SYNTHESIS OF CERIA NANOPlate WITH HYDROGEN PEROXIDE PRECURSOR (TANG, 2005)

| Vol. 30% H₂O₂ | 8 ml |
| Solution 0.05 M Ce(NO₃)₃ | 8 ml |
| Stirring time | 10 min |
| Temperature | 220°C |
| Time | 24 h |

**TABLE III-II**

SYNTHESIS OF CERIA NANOPlate WITH AMMONIUM HYDROXIDE PRECURSOR

| Vol. NH₄OH (28%) | 10 ml |
| Solution 0.05 M Ce(NO₃)₃ | 10 ml |
| Stirring time | 10 min |
| Temperature | 100°C |
| Time | 24 h |
### TABLE III-III

SYNTHESIS OF CERIA NANOROD (C_{NaOH}=9M) WITH SODIUM HYDROXIDE PRECURSOR (MAI, 2005)

<p>| | |</p>
<table>
<thead>
<tr>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>Weight of 0.05 M Ce(NO$_3$)$_3$</td>
<td>0.86 g</td>
</tr>
<tr>
<td>DI water</td>
<td>5 ml</td>
</tr>
<tr>
<td>Weight of NaOH pellets</td>
<td>≈12.68 g</td>
</tr>
<tr>
<td>DI water</td>
<td>35 ml</td>
</tr>
<tr>
<td>Total mixed solution</td>
<td>40 ml</td>
</tr>
<tr>
<td>Stirring time</td>
<td>30 min</td>
</tr>
<tr>
<td>Temperature</td>
<td>100°C</td>
</tr>
<tr>
<td>Time</td>
<td>24 h</td>
</tr>
</tbody>
</table>

### TABLE III-IV

SYNTHESIS OF CERIA NANOCUBE (C_{NaOH}=6M) WITH SODIUM HYDROXIDE PRECURSOR (MAI, 2005)

<p>| | |</p>
<table>
<thead>
<tr>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>Weight of 0.05 M Ce(NO$_3$)$_3$</td>
<td>0.86 g</td>
</tr>
<tr>
<td>DI water</td>
<td>5 ml</td>
</tr>
<tr>
<td>Weight of NaOH pellets</td>
<td>≈8.46 g</td>
</tr>
<tr>
<td>DI water</td>
<td>35 ml</td>
</tr>
<tr>
<td>Total mixed solution</td>
<td>40 ml</td>
</tr>
<tr>
<td>Stirring time</td>
<td>30 min</td>
</tr>
<tr>
<td>Temperature</td>
<td>180°C</td>
</tr>
<tr>
<td>Time</td>
<td>24 h</td>
</tr>
</tbody>
</table>

After the hydrothermal synthesis, the precipitates were separated by centrifugation and washed with deionized water several times. In order to characterize, precipitates were dried overnight at 75°C. The precipitates turned into yellow and white powders for NaOH precursor at 6M and 9M, respectively, white for H$_2$O$_2$ and dark yellow for NH$_4$OH. For XRD, CeO$_2$ powders were calcined at 100°C, 400°C, and 600°C for 12 hours.
D. Characterization of the Prepared CeO₂

The crystallinity of final CeO₂ nanostructures was recorded by XRD (D8 Discover) shown in Figure 3.6 using CuKα radiation (\(\lambda = 1.5406 \text{ Å}\)) in the 2θ range of 20-90° at 0.02° 2θ/s. To estimate crystallite size and lattice parameter, the reflection from (111) plane was used for the analysis. Particle morphology, state of agglomeration and crystalline state were determined by transmission electron microscopy (TEM) shown in Figure 3.4. The Brunauer Emmett Teller (BET) specific surface area was measured by nitrogen-adsorption-desorption isotherms at 77.3 K with TriStar 3000. X-ray photoelectron spectra (XPS) shown in Figure 3.5 were measured by employing Mg-Kα radiation (BE = 1253.6 eV). The binding energy (BE) was calibrated by setting the BE of Ce 3d, O 1S, and C 1S to 935-875, 545-525 and 295-275 eV, respectively.

E. Equipment

![FIGURE 3.2 – Vulcan Furnace](image)

Ney® Vulcan 3-550 Furnace
Dentsupply Ceramco International
York, PA 17404
Mfg No.: AKW0917113
Serial No.: 9493308
FIGURE 3.3 – Hydrothermal Autoclave

FIGURE 3.4 - Transmission Electron Microscopy

Teflon autoclave (50 ml)

Tecnai F20
FEI Company
Made in USA
Thermo VG Scientific  
Product Type: Multilab  
Serial No.: A1126  
The Birches Industrial Estate  
West Sussex, United Kingdom RH91UB

FIGURE 3.5 - X-ray Photoelectron Spectroscopy

Bruker AXS – Diffraktometer D8  
Erz Nr.: 7KP2025 – 2LR15-0Z  
Serial No.: 203407  
Karlsruhe, Germany D76181

FIGURE 3.6 - X-ray diffraction
IV. RESULTS AND DISCUSSION

A. Shape Selective Synthesis and Its Mechanism

Previously, it was stated that ceria nanoplates, nanocubes and nanorods were fabricated by using three different precursors (H$_2$O$_2$, NH$_4$OH, and NaOH) and using Ce(NO$_3$)$_3$ as the cerium source. In this work, we found that CeO$_2$ nanoplates using H$_2$O$_2$ and NH$_4$OH are obtained in the investigated temperature range from 100°C to 250°C. In the case of nanoplates using H$_2$O$_2$, increasing the hydrothermal temperature or decreasing the time does not affect the crystal structure. Furthermore, nanorods and nanocubes could be selectively fabricated by changing the base concentrations of NaOH and hydrothermal temperature. CeO$_2$ nanocubes were synthesized at a hydrothermal temperature of 100°C with a base concentration ($C_{\text{NaOH}} = 6$ M) whereas, the elevated concentration ($C_{\text{NaOH}} = 9$ M) led to the formation of CeO$_2$ nanorods. According to the results discussed later, the base concentration and hydrothermal temperature seem to be the two key factors in the selective formation of CeO$_2$ nanocubes and nanorods.

In the following, a sequence of reaction condition-dependent experiments were conducted to understand the characteristics of the nucleation and crystal growth processes involved in the synthesis and further to uncover the underlying shape-selective mechanism. Details of ceria characterization and analysis of the different morphologies help to understand this phenomenon.
B. TEM analysis

A TEM image of CeO₂ nanoplates using H₂O₂ as oxidizing agent shown in Figure 4.1 has an average particle size of 25 x 20 x 20 nm. The SAED (selected area electron diffraction) pattern shows that the nanoplates were composed of cubic fluorite-type CeO₂ described later. Three kinds of lattice fringe directions attributed to (111), (200), and (220) were observed for these nanoplates, which have respective interplanar spacing of 0.31, 0.27, and 0.19 nm calculated using Bragg’s Law discussed in the section E. XRD analysis.

FIGURE 4.1 - TEM and SAED images of CeO₂ nanoplates at 220°C by H₂O₂.

CeO₂ nanoplates were formed using H₂O₂ at different temperatures but maintaining the same concentration to see if the synthesis temperature has an effect on ceria crystal structure. However, it did not show any difference by synthesizing at temperatures of 180°C and 200°C, for 24 hours. Also, it did not show any effect at 250°C.
by changing the hydrothermal time to 3 hours as shown in Figure 4.2. They all present uniform nanoplate structure which indicate that temperature did not have an effect in the nanoceria structure using H₂O₂.

![Figure 4.2](image)

**FIGURE 4.2** - TEM, HRTEM and SAED images of CeO₂ nanoplates by H₂O₂ synthesized at different temperatures (a) 180°C for 24 hr, (b) 200°C for 24 hr, and (c) 250°C for 3h.

Ceria nanoplates were formed by using NH₄OH as mentioned previously. Figure 4.3 shows ceria nanoplates synthesized at 100°C and agglomeration can be seen. Agglomeration usually increases with decrease in particle size due to higher surface energy (Tok, 2007). Nanoplates synthesized by NH₄OH also exhibit the lattice planes
(111), (200), and (220), same interplanar spacings and has a BET surface area of 62.01 m$^2$/g which will be explained further.

![TEM and SAED image of CeO$_2$ nanoplates by NH$_4$OH at 100°C.]

**FIGURE 4.3** - TEM and SAED image of CeO$_2$ nanoplates by NH$_4$OH at 100°C.

Figure 4.4 depicts the TEM and HRTEM images of CeO$_2$ nanorods using NaOH as the precursor with a concentration of 9 M. Nanorods show a 1D growth structure with a preferred growth direction along [110], and are enclosed by (220) and
(200) planes in other studies (Mai, 2005). In this study, nanorods are enclosed by the lattice planes (111), (200) and (220) according to FFT (fast Fourier transform) analysis, inset in Figure 4.4, with the direction along the [110]. Nanorods exposed the most the (111) plane with an interplanar spacing of 0.31 nm. The nanorods have an average diameter 12-20 nm and a length from 100 to 200 nm.

![TEM and HRTEM images of CeO2 nanorods](image)

**FIGURE 4.4-** TEM and HRTEM images of CeO$_2$ nanorods with a C$_{NaOH}$= 9 M at 100$^\circ$C.

In addition, ceria was synthesized by keeping the same concentration of NaOH (9 M) and changing the hydrothermal temperature to 180$^\circ$C instead of 100$^\circ$C. The crystal structure changes from nanorods to nanocubes. TEM and HRTEM images of ceria nanocubes are shown in Figure 4.5. This result indicates that temperature does have an effect on the crystal structure. The average size of nanocubes is about 40 x 40 x 30 nm. We can see that ceria nanocubes can also be fabricated either using a C$_{NaOH}$ = 9 M at 180$^\circ$C or C$_{NaOH}$= 6 M at 100$^\circ$C.
FIGURE 4.5 - TEM and HRTEM images of CeO$_2$ nanocubes with a $C_{\text{NaOH}} = 9$ M at 180$^\circ$C.

As mentioned earlier, ceria nanocubes are developed by using NaOH with a concentration of 6 M at hydrothermal temperature at 100$^\circ$C. Figure 4.6 illustrates the TEM and HRTEM images of CeO$_2$ nanocubes with an average size of 30 x 30 x 20 nm. Through FFT analysis, nanocubes are enclosed by the (111), (200), and (220) planes. The HRTEM displays the clear (111) lattice fringe in the directions [100] and [011] with an interplanar spacing of 0.31 nm. Other studies imply that CeO$_2$ nanocubes are only enclosed by the (200) planes (Mai, 2005). The TEM and HRTEM images of CeO$_2$ nanoplates, nanorods, and nanocubes in Figures 4.1 to 4.6 show that they are of single crystalline nature.
FIGURE 4.6 - TEM and HTERM images of CeO$_2$ nanocubes with a $C_{NaOH}$ = 6 M at 180°C.

C. BET specific surface area analysis

Brunaure Emmett Teller (BET) analysis was performed to measure the surface area of different CeO$_2$ nanostructures. Table IV-I reports the BET specific surface area of ceria nanoplates, nanocubes and nanorods. Nanoplates show a higher surface area while nanorods and nanocubes have lower surface area. Ceria nanostructures need to have a high surface area because of their importance in industrial applications such as catalyst. In the present work, nanorods and nanocubes show a low surface area, and, as a result, CeO$_2$ nanoplates could be better catalyst from surface area aspect. CeO$_2$ nanoplates have high surface area than nanocubes and nanorods. Nanorods and nanocubes have compatible surface area even though they are of different morphology.
TABLE IV-I

BET SPECIFIC SURFACE AREA OF DIFFERENT CeO$_2$ SHAPES

<table>
<thead>
<tr>
<th>Precursor</th>
<th>H$_2$O$_2$</th>
<th>NH$_4$OH</th>
<th>NaOH-6M</th>
<th>NaOH-9M</th>
</tr>
</thead>
<tbody>
<tr>
<td>Shape</td>
<td>Nanoplate</td>
<td>Nanoplate</td>
<td>Nanocube</td>
<td>Nanorod</td>
</tr>
<tr>
<td>BET surface area (m$^2$/g)</td>
<td>95.11</td>
<td>62.07</td>
<td>16.55</td>
<td>16.74</td>
</tr>
</tbody>
</table>

The N$_2$ adsorption-desorption isotherm plots are illustrated in Figure 4.7. The hysteresis loop at P/P$_0$ 0.8-1.0 of Fig. 4.7c indicates that nanocubes elucidate the typical IUPAC type IV pattern (Pavasupree, 2006), revealing the presence of mesoporous material. The isotherm for nanorods in Fig. 4.7d exhibits also type IV pattern with variation of nitrogen adsorbed volume at P/P$_0$ 0.4-0.9 indicating also the existence of mesopores. However, BET surface area does not correlate with the existence of mesoporous material for nanorods and nanocubes because they show small surface area.
FIGURE 4.7 - Nitrogen adsorption-desorption isotherm pattern as synthesized of CeO$_2$ nanoplates using H$_2$O$_2$ (a), nanoplates using NH$_4$OH (b), nanocubes (c), and nanorods (d), (● adsorption, ■ desorption).

Nanoplates also present a typical type IV pattern with hysteresis loop as shown in Fig. 4.7a-b, which is characteristic of mesoporous material according to the classification of
IUPAC. Characteristic features of the Type IV isotherm are its hysteresis loop, which is associated with capillary condensation taking place in mesopores. Type IV isotherms are given by many mesoporous industrial adsorbents (Sing, 1985). A sharp increase in adsorption volume of N$_2$ was observed and located in the P/P$_o$ range of 0.70-0.85. This sharp increase can be attributed to the capillary condensation, indicating a good homogeneity of the sample and fairly small pore size since the P/P$_o$ position of the inflection point is related to the pore dimension (Pavasupree, 2006). The pore size distribution based on Barret-Joyner-Halenda (BJH) model of CeO$_2$ at different shapes is shown in Figure 4.8.

FIGURE 4.8 - Plot of the pore volume vs. pore diameter for CeO$_2$ nanoplates by H$_2$O$_2$, nanoplates by NH$_4$OH, nanocubes, and nanorods in BJH adsorption (a) and desorption (b).
We can see from Figure 4.8 of adsorption and desorption pore diameter distribution that the application of BJH model to desorption branch will give a completely different result compared to that of the adsorption branch. Similar phenomena have been observed by others (Groen, 2003). In desorption branch, nanocubes and nanorods isotherms have narrower pores centered on 13 nm and 7 nm, respectively. However, in the adsorption branch, nanorods and nanocubes show a broad distribution centered approximately 17 nm and 9 nm, respectively. Nanoplates by NH₄OH derived from the adsorption branch indicates that the created mesopores show a semi-narrow distribution by having an average pore diameter of about 8-9 nm while nanoplates by H₂O₂ had about 7-8 nm. The pore sizes of the CeO₂ nanomaterials are within the mesopores range 2-50 nm (Sing, 1985), which prove that they are of mesoporous material. Table IV-II summarizes the pore diameter and pore volume for the ceria structures studied by BJH adsorption.

<table>
<thead>
<tr>
<th>Precursor</th>
<th>Shape</th>
<th>Pore Volume by BJH Adsorption (cm³/g)</th>
<th>Pore size by BJH Adsorption (nm)</th>
</tr>
</thead>
<tbody>
<tr>
<td>H₂O₂</td>
<td>Nanoplate</td>
<td>0.179</td>
<td>7.33</td>
</tr>
<tr>
<td>NH₄OH</td>
<td>Nanoplate</td>
<td>0.149</td>
<td>8.54</td>
</tr>
<tr>
<td>NaOH-6M</td>
<td>Nanocube</td>
<td>0.073</td>
<td>14.75</td>
</tr>
<tr>
<td>NaOH-9M</td>
<td>Nanorod</td>
<td>0.042</td>
<td>8.52</td>
</tr>
</tbody>
</table>
D. XPS analysis

X-ray photoelectron spectroscopy (XPS) was conducted in the studied nanoceria structures. Figure 4.9 exhibits the XPS patterns of CeO$_2$ nanoplates, nanocubes, and nanorods. Peaks of Ce 3d, O 1s, C 1s, and Ce 4d can be recognized. Also, there were no peaks observed attributable to Na 2p and ammonium which means that the NaOH and NH$_4$OH impurities were not present.

FIGURE 4.9 - XPS survey spectra of CeO$_2$ nanoplates by H$_2$O$_2$ (a), nanocubes (b), nanoplates by NH$_4$OH (c), and nanorods (d).
Moreover, Figure 4.10 indicates that six Ce 3d binding energy (BE) peaks were steady with previous studies on Ce\(^{4+}\) indicating that the main cerium valence of cerium oxide in the nanoplates, nanocubes, and nanorods was Ce\(^{4+}\). Ceria nanoplates using H\(_2\)O\(_2\) show a higher intensity than the rest while nanoplates using NH\(_4\)OH and nanocubes have similar intensities.

![XPS spectrum of the six peaks of Ce 3d for nanoplates by H\(_2\)O\(_2\) (a), nanocubes (b), nanoplates by NH\(_4\)OH (c), and nanorods (d).](image)

The mixed valence state (Ce\(^{3+}\) and Ce\(^{4+}\)) is illustrated in Figure 4.11 for the synthesized cerium oxide nanoplates by H\(_2\)O\(_2\). Ceria nanoplates, nanocubes and nanorods also show the same presence of a mixed valence state. Ce\(^{3+}\) ions are introduced in the nanoceria crystal lattice because of oxygen vacancies created by surface chemical
reactions (Tarnuzzer, 2005). Also, Tarnuzzer suggested that nanoceria behaves as an antioxidant due to the presence of the mixed valence states of Ce$^{3+}$ and Ce$^{4+}$ on the surface.

![XPS spectrum of the synthesized ceria nanoplates by H$_2$O$_2$ with the mixed valence state.](image)

**FIGURE 4.11** - XPS spectrum of the synthesized ceria nanoplates by H$_2$O$_2$ with the mixed valence state.

Deconvoluted XPS Ce (3d) spectra were performed in CeO$_2$ nanoplates, nanocubes, and nanorods. An example image of deconvoluted peaks of ceria nanocubes is shown in Figure 4.12. Similar peak fitting for the XPS spectra of the other CeO$_2$ nanomaterials are presented in Appendix 2. The peaks $v_0$, $v'$, $u_0$, and $u'$ are characteristic of Ce$^{3+}$ while $v''$, $v'''$, $u$, $u''$, and $u'''$ are attributed to Ce$^{4+}$ (Seal, 2001). In the present
deconvoluted spectrums, peak \( \nu \) has a value of zero because it was not necessary to add another peak to fit it into the actual spectrum.

![Deconvoluted XPS Ce (3d) spectrum of ceria nanocubes](image)

FIGURE 4.12 - Deconvoluted XPS Ce (3d) spectrum of ceria nanocubes.

The individual peak positions for all the studied nanoceria shapes are listed in Table IV-III. Notice that the values of the individual peaks are essentially identical between the different shapes of ceria which indicates that the fitted-peaks were introduced in similar positions in order to fit the actual Ce 3d spectrum.
**TABLE IV-III**

**XPS BINDING ENERGIES OF INDIVIDUAL PEAKS OF THE CE (3D) SPECTRUM**

**FOR DIFFERENT CeO₂ SHAPES**

<table>
<thead>
<tr>
<th>Shape</th>
<th>Ce (3d₅/₂)</th>
<th></th>
<th></th>
<th>Ce (3d₃/₂)</th>
<th></th>
<th></th>
<th></th>
<th></th>
<th></th>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>v₀</td>
<td>v'</td>
<td>v''</td>
<td>v'''</td>
<td>u₀</td>
<td>u</td>
<td>u'</td>
<td>u''</td>
<td>u'''</td>
<td></td>
<td></td>
</tr>
<tr>
<td>Nanoplate (H₂O₂)</td>
<td>882.32</td>
<td>885.17</td>
<td>888.39</td>
<td>891.00</td>
<td>898.40</td>
<td>901.07</td>
<td>903.61</td>
<td>907.07</td>
<td>916.72</td>
<td></td>
<td></td>
</tr>
<tr>
<td>Nanoplate (NH₄OH)</td>
<td>882.14</td>
<td>884.26</td>
<td>888.27</td>
<td>890.81</td>
<td>898.28</td>
<td>900.88</td>
<td>903.55</td>
<td>907.56</td>
<td>916.62</td>
<td></td>
<td></td>
</tr>
<tr>
<td>Nanocube</td>
<td>881.78</td>
<td>884.99</td>
<td>887.72</td>
<td>890.27</td>
<td>898.03</td>
<td>900.70</td>
<td>903.37</td>
<td>907.31</td>
<td>916.23</td>
<td></td>
<td></td>
</tr>
<tr>
<td>Nanorod</td>
<td>882.08</td>
<td>885.29</td>
<td>887.78</td>
<td>890.33</td>
<td>897.97</td>
<td>900.70</td>
<td>903.07</td>
<td>906.77</td>
<td>916.05</td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

A partial quantitative analysis of the integrated peak area can provide the concentration of Ce³⁺ ions in the synthesized ceria. Equation 1 calculates the fraction of Ce³⁺ (Seal, 2001):

\[
[\text{Ce}^{3+}] = \frac{A_{v_0} + A_{v'} + A_{u_0} + A_{u'}}{A_{v_0} + A_{v'} + A_{u_0} + A_{u'} + A_{v''} + A_{v'''} + A_{u''} + A_{u'''}} (1)
\]

where \( A_i \) is the integrated area of peak “i.” Table IV-IV and IV-V present the integrated areas obtained from each peak and the concentration of Ce³⁺ of ceria nanoplates, nanocubes and nanorods, respectively.
Using Equation 1, we can calculate the concentration of Ce$^{3+}$. It has been reported that the concentration of Ce$^{3+}$ increases with a decrease in particle size of nanoceria (Deshpande, 2005). In the current study, we found that nanocubes and nanorods have high concentration of Ce$^{3+}$ and smaller particle sizes. It can be attributed to higher oxygen vacancies in the smaller particles.
E. XRD analysis

The crystallite size and crystal structure of the CeO$_2$ nanoplates, nanocubes and nanorods are obtained by examining the XRD patterns. Figure 4.13 exhibits the XRD patterns of the CeO$_2$ nanomaterials after calcined at 100$^\circ$C. The exposing XRD peaks correspond to the (111), (200), (220), (311), (400), (3310), (420), and (422) planes of a ceria cubic fluorite structure (space group Fm3m (225)) with an average lattice parameter $a = 5.407$ Å. The XRD peaks for nanorods (c) and nanoplates using (NH$_4$OH) (d) are very broad which indicates that these two morphologies have low crystallinity or contain a large portion of amorphous phase (Sujama, 2008).

![XRD patterns of CeO$_2$ nanomaterials](image)

FIGURE 4.13 - XRD of CeO$_2$ nanoplates by H$_2$O$_2$ (a), nanocubes (b), nanoplates by NH$_4$OH (c), and nanorods (d).
The XRD patterns for CeO$_2$ nanoplates, nanocubes and nanorods calcined at 400°C and 600°C are illustrated in Figures 4.14 to 4.17. The nanoplates exhibit in Figure 4.14 using H$_2$O$_2$ show a higher intensity at 100°C than at 400°C while nanoplates using NH$_4$OH shown in Figure 4.15 do not present any difference between 100°C and 600°C. XRD patterns for nanocubes (C$_{NaOH}$=6M) shown in Figure 4.16 at 100°C shows small peaks in between the (200) plane, either another crystal structure is present or some NaOH impurities. However, at 400°C, the (200) plane shows a uniform peak which denies the fact of another crystal structure leaving the possibility of some residues of NaOH impurities in the sample. For nanorods, XRD spectrum is wider at 100°C and has lower intensity than at 600°C as shown in Figure 4.17. Narrow peaks show high crystallinity while wide peaks, low crystallinity.

FIGURE 4.14 - XRD patterns of CeO$_2$ nanoplates by H$_2$O$_2$ calcined at 100°C (a) and 400°C (b).
FIGURE 4.15 - XRD patterns of CeO$_2$ nanoplates by NH$_4$OH calcined at 100$^\circ$C (a) and 600$^\circ$C (b).

FIGURE 4.16 - XRD patterns of CeO$_2$ nanocubes (C$_{\text{NaOH}}$=6M) calcined at 100$^\circ$C (a) and 400$^\circ$C (b).
Table IV-VI shows the calculated lattice parameters, \( a \), using Equation 2 versus the theoretical XRD values at different calcinations temperatures.

\[
a = d_{hkl} \sqrt{h^2 + k^2 + l^2}
\]  

(2)

The experimental lattice parameters are similar to the theoretical ones proving that the diffraction peaks correspond only to CeO\(_2\). The different lattice parameters values are due to the different precursors used to synthesize CeO\(_2\). It was necessary to find the right ceria pattern that belongs to ceria with its respective precursor. It is noted that some values of the lattice parameter for CeO\(_2\) nanostructures are lower or higher than that reported for the bulk of CeO\(_2\) which is \( a = 5.411 \) Å provided in other studies (Zhou,
2005). The lattice parameter only varied within the average range ($\Delta a/a \approx 0.25\%$). In addition, the (111) plane is the most exposing plane in the XRD spectra.

### TABLE IV-VI

**COMPARISON OF LATTICE PARAMETERS AT DIFFERENT CALCINATION TEMPERATURES**

<table>
<thead>
<tr>
<th>Precursor</th>
<th>Shape</th>
<th>100$^\circ$C</th>
<th>400$^\circ$C</th>
<th>600$^\circ$C</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>$a_0$, calculated</td>
<td>$a_0$, from XRD</td>
<td>$a_0$, calculated</td>
<td>$a_0$, from XRD</td>
</tr>
<tr>
<td>H$_2$O$_2$ Nanoplate</td>
<td>5.4291</td>
<td>5.4124</td>
<td>5.4382</td>
<td>5.4124</td>
</tr>
<tr>
<td>NaOH Nanocube</td>
<td>5.4035</td>
<td>5.4037</td>
<td>5.4344</td>
<td>5.4124</td>
</tr>
<tr>
<td>NaOH Nanorod</td>
<td>5.4157</td>
<td>5.4037</td>
<td>-</td>
<td>-</td>
</tr>
<tr>
<td>NH$_4$OH Nanoplate</td>
<td>5.4291</td>
<td>5.4037</td>
<td>-</td>
<td>-</td>
</tr>
</tbody>
</table>

In order to calculate the crystallite size ($D_{hkl}$), the following equations were needed:

\[ n\lambda = 2d_{hkl}\sin \theta \quad (3) \]

\[ D_{hkl} = \frac{0.89 \times \lambda}{B_{hkl} \times \cos \theta_B} \quad (4) \]

Equation 3 expresses Bragg’s Law where $n = 1$, $\lambda = 1.5406$ Å and $\theta$ is obtained from XRD spectra. The $a$ refers to the CeO$_2$ FCC (face-centered-cubic) lattice parameter, and $h,k,l$ are the crystallite size indices while $d$ is the crystalline face space. On the other hand, Equation 4 is the Scherrer equation which is used to calculate the crystallite size ($D_{hkl}$) (Tennyson, 2007). The $B_{hkl}$ corresponds to the measured half-width at maximum peak height. Table IV-VII summarizes the lattice type, lattice parameter and crystallite...
size values of CeO₂ nanostructures by using the equations mentioned earlier. The crystallite sizes for nanoplates, nanocubes and nanorods were estimated from X-ray line broadening of the reflections of (111) plane.

**TABLE IV-VII**

**LATTICE TYPE AND CRYSTALLITE SIZE OF CeO₂ NANOSTRUCTURES**

<table>
<thead>
<tr>
<th>Shape</th>
<th>As synthesized</th>
<th>Calcined at 400°C</th>
<th>Calcined at 600°C</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>Lattice Type</td>
<td>$a$ (Å)</td>
<td>$D_{hkl}$ (nm)</td>
</tr>
<tr>
<td>Nanoplate (H₂O₂)</td>
<td>Cubic</td>
<td>5.4291</td>
<td>24.92</td>
</tr>
<tr>
<td>Nanocube</td>
<td>Cubic</td>
<td>5.4035</td>
<td>19.65</td>
</tr>
<tr>
<td>Nanorod</td>
<td>Cubic</td>
<td>5.4157</td>
<td>5.97</td>
</tr>
<tr>
<td>Nanoplate (NH₄OH)</td>
<td>Cubic</td>
<td>5.4291</td>
<td>12.30</td>
</tr>
</tbody>
</table>

The lattice type is cubic for the different morphologies of ceria. Lattice parameters values are slightly different to each other except for nanoplates by NH₄OH which are the same regardless of the calcination temperatures. Figure 4.18 illustrates the effect of calcination temperature and lattice parameter on the crystallite size. We can see that crystallite size in nanoplates using H₂O₂ shown in Fig. 4.18a increases at high calcination temperature and lattice parameter while nanoplates using NH₄OH in Fig 4.18c remain unchanged or steady, indicating that temperature and lattice parameter do not affect the size. Nanocube’s crystallite size in Fig 4.18b decreases at high calcination temperature and lattice parameter while the crystallite size of nanorods in Fig 4.18d is proportional to temperature and lattice parameter.
Figure 4.18 - Crystallite size as a function of calcination temperatures for nanoplates by H$_2$O$_2$ (a), nanocubes (b), nanoplates by NH$_4$OH (c), and nanorods (d).

By comparing the crystallite size and surface area of ceria nanorods, nanoplates, and nanocubes: nanoplates have high surface area with respect to their crystallite size while nanocubes exhibit high crystallite size at low surface area.
Theoretically, BET specific surface area should increase with the decrease in crystallite size (Tok, 2007). However, this theory only applies to nanorods and nanoplates. In the case of nanocubes, low BET surface area can be accounted for the effect of agglomeration. The agglomerations of these morphologies can be observed in the TEM images described earlier. Table IV-VIII presents the comparison of the surface area with the crystallite size. On the other hand, Figure 4.19 illustrates the concentration of Ce$^{3+}$ as a function of lattice parameter. Higher values of Ce$^{3+}$ ions lead to a larger lattice strain and, consequently, an increase in the lattice parameter of the particles (Deshpande, 2005). Though, in this study, high concentration of Ce$^{3+}$ shows low lattice parameter. This is due to high oxygen vacancies which can induce the lattice parameter to change.

TABLE IV-VIII

BET SPECIFIC SURFACE AREA AND CYRSTALLITE SIZE OF DIFFERENT CeO$_2$ SHAPES

<table>
<thead>
<tr>
<th>Precursor</th>
<th>H$_2$O$_2$</th>
<th>NH$_4$OH</th>
<th>NaOH-6M</th>
<th>NaOH-9M</th>
</tr>
</thead>
<tbody>
<tr>
<td>Shape</td>
<td>Nanoplate</td>
<td>Nanoplate</td>
<td>Nanocube</td>
<td>Nanorod</td>
</tr>
<tr>
<td>BET surface area (m$^2$/g)</td>
<td>95.11</td>
<td>62.07</td>
<td>16.55</td>
<td>16.74</td>
</tr>
<tr>
<td>Crystallite size (nm)</td>
<td>24.92</td>
<td>12.30</td>
<td>19.65</td>
<td>5.97</td>
</tr>
</tbody>
</table>
In summary, from all the characterizations performed to the synthesized CeO$_2$ nanomaterials, we can conclude that the precursors and reaction parameters have a significant effect on the morphology of ceria. Nanoplates, nanorods and nanocubes presented different and unique properties. To enhance catalytic performance, CO oxidation is an important process in the three-way catalysis. In order to design and synthesize CeO$_2$ nanostructured catalyst, decrease in the less reactive crystal planes and increase in the more reactive ones is needed to optimize the desired structure of the active sites. However, it has been a very difficult task for most of the practical catalysts (Kung, 2003). In general, high-surface area nanocatalytic materials exhibiting numerous crystal faces, edges, and corners, which are conventionally considered active sites for the
adsorption of reactants (Terribile, 1997; Daturi, 2000 and Stoimenov, 2003), should generate better catalytic performance.

Thus, it is reasonable to consider that the different catalytic properties of various CeO$_2$ nanomaterials may occur from their morphologies and sizes. The CO oxidation activity of CeO$_2$ is strongly influenced by the crystal planes due to their different ability of creating oxygen vacancies which are the key point for the CO oxidation. In the present study, the frequently exposed crystal plane in CeO$_2$ nanoplates, nanorods and nanocubes is (111) plane. The (111) is the most stable surface which has the lowest surface energy (Lundberg, 2004), followed by (110), (100), and (211). Sayle et al. (2002) found that (100) is more reactive and important catalytically as compared with either (110) or the (111) plane. Previous computer modeling demonstrates that less energy is required to form vacancies on (100) than that on (110) and (111). This results suggest that the exposure of more reactive (100) surfaces should facilitate the formation of oxygen vacancies and enhance the OSC (oxygen storage capacity) of CeO$_2$ (Mai, 2005). Lundberg et al. (2004) found that the more exposure of (100) surfaces in ceria can greatly enhance its CO activity.

This research shows that ceria nanoplates synthesized with H$_2$O$_2$ and NH$_4$OH are mesoporous and have higher surface area than nanocubes and nanorods. Nanorods have compatible surface area as nanocubes synthesized with NaOH, however, they have more exposed crystalline planes (111), (200), and (220) than nanocubes and nanoplates. To provide active sites for catalytic reactions, it is necessary that the exposed planes with higher surface energy could be generated and stabilized (Zhou, 2005). We found that nanorods have also higher concentration of Ce$^{3+}$ than nanoplates which is
important for OSC and redox reactions. Nanorods also have smaller crystallite size. The XPS, BET and XRD results in this thesis elucidate that nanorods are better catalyst for CO oxidation and OSC (Mai 2005, Zhou 2005, Tana 2009).
V. CONCLUSIONS

The selection of precursors (H₂O₂, NaOH and NH₄OH) in the synthesis of ceria has yielded different results in terms of crystallite size, crystallinity, particle size, and specific surface area and oxygen vacancy. CeO₂ nanoplates, nanocubes, and nanorods were studied by TEM, BET specific surface area, XPS and XRD. Nanoplates (H₂O₂) and nanorods had a crystallite size range of 22-25nm and 5-11 nm, respectively. The crystallite size was proportional to the calcination temperature and lattice parameter. Nanocubes’ crystallite size decreased with the increase of calcination temperature and lattice parameter while nanoplates (NH₄OH) remained constant in the size of 12.3 nm regardless of temperature and lattice parameter. From TEM, agglomeration is observed in nanoplates. Mesoporous material was found in nanoplates by H₂O₂ and NH₄OH, nanorods and nanocubes due to their adsorption-desorption isotherm and pore size distribution. Their pore size was in the range of 7-15 nm according to BJH adsorption model. CeO₂ structures share the same exposing plane (111) and same interplanar d-spacings although their lattice parameters were slightly different from each other. Nanoplates with H₂O₂ have larger particle size than nanoplates with NH₄OH; nanorods have smaller particle size than nanocubes. After the characterization and analysis, nanorods are considered better catalyst for CO oxidation and oxygen storage capacity (OSC).
VI. RECOMMENDATIONS

It was found that high temperature did not affect the crystal structure of ceria nanoplates synthesized using H$_2$O$_2$, so it is necessary to conduct synthesis of nanoplates at different concentrations of H$_2$O$_2$ and observe any significant differences in size and crystal structure. Ceria nanoplates using NH$_4$OH need to be studied at higher synthesis temperatures and concentrations of NH$_4$OH since is uncertain the effect of higher temperature and concentration on the shape and size of nanoplates. For nanorods and nanocubes, decrease the hydrothermal time to see if the crystal size and structure change. Also, it will be helpful if the precipitates are stored longer before characterization and note if time has a significant effect on the crystal structure of ceria.

As mentioned in other studies, other types of characterizations such as particle size analyzer, FT-IR, and TG-DTA will be helpful to better understand the morphology and particle size of ceria nanostructures. Finally, it is necessary to perform CO oxidation to find out the catalytic evaluation and oxygen storage capacity by H$_2$- TPR in order to understand this behavior. TEM images of ceria nanomaterials after calcination are needed to evaluate their exposing planes.
APPENDIX 1–RAW DATA

TABLE A1 – 1
EXPERIMENTAL DATA FOR CERIA NANOPLATES USING H2O2

<table>
<thead>
<tr>
<th>Conc. (ml) [0.05M Ce(NO₃)₃]</th>
<th>Conc. 30% H₂O₂ (ml)</th>
<th>Color (Before)</th>
<th>Temperature (°C)</th>
<th>Time (hr)</th>
<th>Color (After)</th>
</tr>
</thead>
<tbody>
<tr>
<td>8</td>
<td>8</td>
<td>orangish</td>
<td>220</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>8</td>
<td>8</td>
<td>orangish</td>
<td>180</td>
<td>24</td>
<td>milky color (white)</td>
</tr>
<tr>
<td>8</td>
<td>10</td>
<td>orangish</td>
<td>100</td>
<td>24</td>
<td>white transparent</td>
</tr>
<tr>
<td>8</td>
<td>8</td>
<td>orangish</td>
<td>200</td>
<td>3</td>
<td>milky color (white)</td>
</tr>
<tr>
<td>8</td>
<td>8</td>
<td>orangish</td>
<td>100</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>8</td>
<td>8</td>
<td>orangish</td>
<td>220</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>8</td>
<td>8</td>
<td>orangish</td>
<td>220</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>8</td>
<td>10</td>
<td>orangish</td>
<td>220</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>8</td>
<td>8</td>
<td>orangish</td>
<td>220</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>8</td>
<td>8</td>
<td>orangish</td>
<td>250</td>
<td>3</td>
<td>white</td>
</tr>
<tr>
<td>8</td>
<td>8</td>
<td>orangish</td>
<td>250</td>
<td>12</td>
<td>white</td>
</tr>
<tr>
<td>8</td>
<td>10</td>
<td>orangish</td>
<td>200</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>8</td>
<td>8</td>
<td>orangish</td>
<td>220</td>
<td>24</td>
<td>white</td>
</tr>
</tbody>
</table>

TABLE A1 – 2
EXPERIMENTAL DATA FOR CERIA NANOPLATES USING NH₄OH

<table>
<thead>
<tr>
<th>Solution [0.05M Ce(NO₃)₃] (ml)</th>
<th>NH₄OH (ml)</th>
<th>Color (Before)</th>
<th>Temperature (°C)</th>
<th>Time (hr)</th>
<th>Color (After)</th>
</tr>
</thead>
<tbody>
<tr>
<td>10</td>
<td>10</td>
<td>dark orange</td>
<td>100</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>10</td>
<td>10</td>
<td>dark orange</td>
<td>150</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>10</td>
<td>10</td>
<td>dark orange</td>
<td>150</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>10</td>
<td>10</td>
<td>dark orange</td>
<td>100</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>10</td>
<td>10</td>
<td>dark orange</td>
<td>100</td>
<td>24</td>
<td>white</td>
</tr>
</tbody>
</table>
### TABLE A1 – 3

EXPERIMENTAL DATA FOR CERIA NANORODS AND NANOCUBES

<table>
<thead>
<tr>
<th>Weight (g) [0.05M Ce(NO₃)₃]</th>
<th>DI Water (ml)</th>
<th>Weight NaOH (g)</th>
<th>DI Water (ml)</th>
<th>Conc. Ce(NO₃)₃ (mol/L)</th>
<th>C₉H₄ (M)</th>
<th>Color (Before)</th>
<th>T (°C)</th>
<th>Time (hr)</th>
<th>Color (After)</th>
</tr>
</thead>
<tbody>
<tr>
<td>0.8682</td>
<td>5</td>
<td>0.0123</td>
<td>35</td>
<td>0.40</td>
<td>0.01</td>
<td>milky</td>
<td>100</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>0.8684</td>
<td>5</td>
<td>12.6389</td>
<td>35</td>
<td>0.40</td>
<td>9.03</td>
<td>milky</td>
<td>100</td>
<td>24</td>
<td>milky</td>
</tr>
<tr>
<td>0.8686</td>
<td>5</td>
<td>5.6823</td>
<td>35</td>
<td>0.40</td>
<td>4.06</td>
<td>milky</td>
<td>100</td>
<td>24</td>
<td>purple slurry</td>
</tr>
<tr>
<td>0.8684</td>
<td>5</td>
<td>2.4543</td>
<td>35</td>
<td>0.40</td>
<td>1.75</td>
<td>milky</td>
<td>100</td>
<td>24</td>
<td>purple slurry</td>
</tr>
<tr>
<td>0.868</td>
<td>5</td>
<td>0.0906</td>
<td>35</td>
<td>0.40</td>
<td>0.06</td>
<td>milky white</td>
<td>100</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>0.8688</td>
<td>5</td>
<td>12.5983</td>
<td>35</td>
<td>0.40</td>
<td>9.00</td>
<td>milky</td>
<td>100</td>
<td>24</td>
<td>cloudy</td>
</tr>
<tr>
<td>0.868</td>
<td>5</td>
<td>0.0143</td>
<td>35</td>
<td>0.40</td>
<td>0.01</td>
<td>white</td>
<td>100</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>0.868</td>
<td>5</td>
<td>8.4255</td>
<td>35</td>
<td>0.40</td>
<td>6.02</td>
<td>milky</td>
<td>100</td>
<td>24</td>
<td>cloudy</td>
</tr>
<tr>
<td>0.8683</td>
<td>5</td>
<td>2.8075</td>
<td>35</td>
<td>0.40</td>
<td>2.01</td>
<td>milky</td>
<td>100</td>
<td>24</td>
<td>purple slurry</td>
</tr>
<tr>
<td>0.8701</td>
<td>5</td>
<td>12.6852</td>
<td>35</td>
<td>0.40</td>
<td>9.06</td>
<td>milky</td>
<td>100</td>
<td>24</td>
<td>cloudy</td>
</tr>
<tr>
<td>0.8686</td>
<td>5</td>
<td>12.6855</td>
<td>35</td>
<td>0.40</td>
<td>9.06</td>
<td>milky</td>
<td>180</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>0.8685</td>
<td>5</td>
<td>5.6822</td>
<td>35</td>
<td>0.40</td>
<td>4.06</td>
<td>milky</td>
<td>100</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>0.8689</td>
<td>5</td>
<td>8.4586</td>
<td>35</td>
<td>0.40</td>
<td>6.04</td>
<td>milky</td>
<td>180</td>
<td>24</td>
<td>white</td>
</tr>
<tr>
<td>0.8686</td>
<td>5</td>
<td>12.6842</td>
<td>35</td>
<td>0.40</td>
<td>9.06</td>
<td>milky</td>
<td>100</td>
<td>24</td>
<td>white</td>
</tr>
</tbody>
</table>
APPENDIX 2 – DECONVOLUTED XPS SPECTRUMS

Peak and integrated area values are presented in Tables IV-III and IV-IV.

FIGURE A2-1 – CeO₂ nanoplates by H₂O₂ deconvoluted XPS spectrum

FIGURE A2-2 – CeO₂ nanoplates by NH₄OH deconvoluted XPS spectrum
FIGURE A2-3 – CeO$_2$ nanorods deconvoluted XPS spectrum

FIGURE A2-4 - CeO$_2$ nanocubes deconvoluted XPS spectrum
APPENDIX 3: SAMPLE CALCULATION

The following sample calculation is for determining the concentration of Ce$^{3+}$ in a nanoplate. The integrated peak area values are presented in Table IV-IV.

\[
[\text{Ce}^{3+}] = \frac{14680+11040+10500+9474.21}{14680+11040+8354.20+15220+10500+6139.87+9474.21+8199.964743.22} = 0.52 \text{ M}
\]

Another sample calculation is for determining the crystallite size and lattice parameter in nanoplates using H$_2$O$_2$ at 100$^\circ$C. Figure 4.14 will be needed for the calculation.

**Data Needed**

\[
\begin{align*}
\lambda &= 1.5406 \, \text{Å} = 0.15406 \, \text{nm} \\
K &= 0.89 \\
(h,k,l) &= (111) \\
2\theta_{\text{High}} &= 28.61 \\
2\theta_{\text{Low}} &= 28.28 \\
2\theta_{\text{B}} &= 28.45
\end{align*}
\]

**Calculate B**

\[
B = (2\theta_{\text{High}} - 2\theta_{\text{Low}}) \times \frac{\pi}{180} = (28.61 - 28.28) \times \frac{\pi}{180} = 0.00568 \, \text{degree}
\]

**Calculate \( \theta_B \) (degree)**

\[
\theta_B = 28.45 \times \frac{\pi}{180} = 0.25 \, \text{(degree)}
\]

**Calculate \( d_{hkl} \)**

\[
\begin{align*}
d_{hkl} &= \frac{1.5406}{2 \times \sin 0.25} = 3.31 \, \text{Å} = 0.31 \, \text{nm}
\end{align*}
\]

**Calculate \( D_{hkl} \)**

\[
\begin{align*}
D_{hkl} &= \frac{0.89 \times 0.15406}{0.00568 \times \cos 0.25} = 24.92 \, \text{nm}
\end{align*}
\]

**Calculate \( a \)**

\[
\begin{align*}
a &= 3.13 \times \sqrt{1^2 + 1^2 + 1^2} = 5.4291 \, \text{Å}
\end{align*}
\]
REFERENCES


VITA

NITZIA CHEONG NG

Chemical Engineering Department
J. B. Speed School of Engineering
University of Louisville
Louisville, KY 40292

EDUCATION

Master of Engineering in Chemical Engineering with Environmental Engineering Certificate, University of Louisville, Louisville, KY, 2010

Bachelor of Science in Chemical Engineering, University of Louisville, Louisville, KY, 2009

High School Diploma, Instituto Panamericano, Panama City, Panama, 2003

HONORS AND AWARDS

Awarded international student scholarship from International Center at University of Louisville for Bachelor of Science degree, 2008

Awarded student section award from Society of Women Engineers, University of Louisville, 2009